KEY – NET IONIC EQUATION PRACTICE (#4)

1) Unbalanced equation: Na_2SO_4 (aq) + $Ba(NO_3)_2$ (aq) → $NaNO_3$ (aq) + $BaSO_4$ (s)
Balanced net ionic equation: $(SO_4)^{-2}$ (aq) + Ba^{+2} (aq) → $BaSO_4$ (s)

2) Unbalanced equation: SrO (aq) + $Al(ClO_4)_3$ (aq) → Sr $(ClO_4)_2$ (aq) + Al_2O_3 (s)
Balanced net ionic equation: $\underline{3}O^{-2}$ (aq) + $\underline{2}Al^{+3}$ (aq) → Al_2O_3 (s)

3) (can write H_2O as HOH if you want to show the individual ions in it)
Unbalanced equation: HNO_3(aq) + KOH (aq) → HOH (l) + KNO_3 (aq)
Balanced net ionic equation: H^{+1} (aq) + OH^{-1} (aq) → HOH (l)

4) Unbalanced equation: $FeBr_3$ (aq) + $AgNO_3$ (aq) → $Fe(NO_3)_3$ (aq) + AgBr (s)
Balanced net ionic equation: Br^{-1} (aq) + Ag^{+1} (aq) → AgBr (s)

5) Unbalanced equation: K_2S (aq) + $Ca(C_2H_3O_2)_2$ (aq) → $KC_2H_3O_2$ (aq) + CaS (s)
Balanced net ionic equation: S^{-2} (aq) + Ca^{+2} (aq) → CaS (s)

6) (can write H_2O as HOH if you want to show the individual ions in it)
Unbalanced equation: HCl (aq) + NH_4OH (aq) → HOH (l) + NH_4Cl (aq)
Balanced net ionic equation: H^{+1} (aq) + OH^{-1} (aq) → HOH (l)

7) Unbalanced equation: $Cu(ClO_3)_2$ (aq) + $(NH_4)_3PO_4$ (aq) → $Cu_3(PO_4)_2$ (s) + NH_4ClO_3 (aq)
Balanced net ionic equation: $\underline{3}Cu^{+2}$ (aq) + $\underline{2}(PO_4)^{-3}$ (aq) → $Cu_3(PO_4)_2$ (s)

8) Unbalanced equation: $Fe(C_2H_3O_2)_2$ (aq) + Na_3PO_4 (aq) → $Fe_3(PO_4)_2$ (s) + $NaC_2H_3O_2$ (aq)
Balanced net ionic equation: $\underline{3}Fe^{+2}$ (aq) + $\underline{2}(PO_4)^{-3}$ (aq) → $Fe_3(PO_4)_2$ (s)

9) (can write H_2O as HOH if you want to show the individual ions in it)
Unbalanced equation: H_3PO_4 (aq) + NaOH (aq) → HOH (l) + Na_3PO_4 (aq)
Balanced net ionic equation: H^{+1} (aq) + OH^{-1} (aq) → HOH (l)

10) Unbalanced equation: $MgCl_2$ (aq) + Li_2CO_3 (aq) → LiCl (aq) + $MgCO_3$ (s)
Balanced net ionic equation: Mg^{+2} (aq) + $(CO_3)^{-2}$ (aq) → $MgCO_3$ (s)